

# Graphite Oxide-TiO<sub>2</sub> Nanocomposite Type Photocatalyst for Methanol Photocatalytic Reforming Reaction

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## **Graphite oxide-TiO<sub>2</sub> nanocomposite type photocatalyst for methanol photocatalytic reforming reaction**

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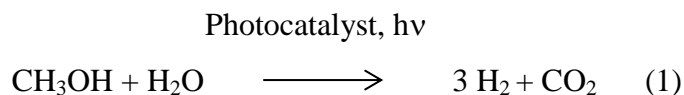
**Abstract:**

Graphite-oxide/TiO<sub>2</sub> (GO/TiO<sub>2</sub>) composite materials were prepared by heterocoagulation method from Brodie's graphite-oxide (GO) in order to test them as catalysts in the methanol photocatalytic reforming reaction in liquid phase. The preparation of the composite itself resulted in only little changes in the structure of GO as it was indicated by attenuated total reflection infrared (ATR-IR) and <sup>13</sup>C magic-angle spinning nuclear magnetic resonance (<sup>13</sup>C MAS NMR) spectroscopic measurements. However, during the photocatalytic reaction, all of the GO/TiO<sub>2</sub> samples darkened strongly indicating structural changes of GO. X-ray photoelectron spectroscopy (XPS) along with NMR confirmed the loss of oxygen functionalities and emergence of graphitic species in the samples recovered from the photocatalytic reaction. Model experiments were designed to identify the key factors determining the activity of the GO/TiO<sub>2</sub> derived photocatalysts. It was found that the emergence of a pronounced coupling between TiO<sub>2</sub> and the graphite-like carbonaceous material is the most important contribution to get active and stable photocatalysts.

**Keywords:** graphite oxide, TiO<sub>2</sub>, methanol, hydrogen, XPS, MAS NMR

## 1. Introduction

Nowadays, hydrogen is regarded as one of the emerging sources of clean energy [1-3]. Photocatalytic hydrogen production is a promising approach for transforming solar energy into chemical energy for storage and transport. Methanol photocatalytic reforming (1) is a potential reaction for large-scale H<sub>2</sub> production [4].



Since Fujishima's and Honda's pioneering work in 1972 [5], there has been a continuously growing interest in TiO<sub>2</sub> as a photocatalyst because of its efficiency, long term stability, cheapness and low toxicity [6-9]. Although TiO<sub>2</sub> is a good candidate for photocatalyst in general, especially in degradation of organic pollutants both in gas phase [7, 10] and in liquid phase [7, 11-13] bare TiO<sub>2</sub> (in the absence of metal co-catalyst) results only in poor H<sub>2</sub> evolution in reaction (1) [14,15]. The activity of TiO<sub>2</sub> in the hydrogen producing reaction can be increased at least by an order of magnitude when a proper co-catalyst is involved [16-18]. The advantages of the co-catalysts are attributed to the suppressed charge recombination; promoted charge separation and transport driven by junctions/interfaces [19-21]. Pt is one of the most effective co-catalysts for H<sub>2</sub> production [16, 20] however it is rather expensive and is available in limited quantities.

Combination of carbon materials with a semiconductor has been reported to enhance the photocatalytic activity [22, 23]. The enhancement mechanism can be related to carbon materials, which (i) offer more active and larger number of adsorption sites and photocatalytic reaction centers, (ii) contribute to the suppression of the recombination of the photogenerated electron/hole pairs, (iii) prolong the lifetime of electrons and holes, (iv) narrow the band gap of photocatalyst, and (v) act as photosensitizer for catalytic reaction [24-25].

Graphite oxide (GO) is an excellent supporting matrix in nanocomposite materials due to its high specific surface area. GO is a layer-structured graphite compound which is relatively cheap, easily available and suitable for mass production. It can be prepared from graphite by electrochemical oxidation [26], most frequently by strong oxidizing agents according to the two main preparation procedures and their more elaborated versions (*e.g.* NaClO<sub>3</sub>/HNO<sub>3</sub> in Brodie's method [27]; KMnO<sub>4</sub>, H<sub>2</sub>SO<sub>4</sub>, NaNO<sub>3</sub> in Hummers-Offeman method [28]). GO is a non-stoichiometric compound with an empirical chemical formula of C<sub>4</sub>O<sub>2</sub>H for well-oxidized samples [29], although the structure of GO depends on the type of the oxidation method [30-31]. Dynamic changes of its structure have also been suggested [32]. GO has weakly acidic character [33] owing to O and H containing functional groups

which are covalently attached to the carbon skeleton [29, 34]. Typical functional groups of GO are -OH, cyclic ether, C=O and COOH. Nevertheless, non-oxidized aromatic regions and/or isolated C=C double bonds are also characteristic for the structure of GO. However, O-containing functional groups provide relatively hydrophilic character of GO [35] so unlike graphene it is easy to use in wet-chemistry [36]. The GO is built up from hydrophilic, stacked graphene-based sheets, exfoliated GO is often added to TiO<sub>2</sub> in aqueous dispersions [37-38] or used as a “base board” for TiO<sub>2</sub> preparation from different Ti precursors [39-41] in order to get reduced graphene oxide (R-GO) containing composites after reduction processes [42-45]. Hydrothermal treatment of the GO-TiO<sub>2</sub> (*i.e.* treatment with the mixture of water and ethanol at elevated temperature in autoclave) also leads to the reduction of the GO, with the formation of R-GO/TiO<sub>2</sub> composite materials [38, 46]. During the hydrothermal treatment removal of O containing functional groups of GO has been observed [47] which has led to re-establishment of conjugated graphene network [48] along with the possible appearance of C-O-Ti bonds [46-47, 49-50]. Composites prepared from GO and TiO<sub>2</sub> have been found beneficial for hydrogen producing reactions [42, 49]. Several works have reported a synergism between reduced graphene oxide (R-GO), TiO<sub>2</sub> and Pt [26, 51].

The key point to taking the benefits of GO is in its exfoliation. Different methods, such as sonication, thermal treatment, microwave treatment, etc., can be used for the above purpose [36]. Sonication seems to be a clean and comfortable method to separate the GO sheets from each other in case of Hummers' GO, however sonication itself is not satisfactory for exfoliation of Brodie's GO [25, 30] and slightly basic conditions are needed to achieve larger degrees of dispersion [33]. However, the application of graphite oxide samples synthesized via Brodie protocol may be superior to Hummers' GO in terms of purity (metal and sulfur-containing species are absent in Brodie-GO which can have a detrimental effect related to the activity of catalytic reactions) [30]. Recently, GO/TiO<sub>2</sub> exfoliated nanocomposite has been prepared from Brodie's GO by heterocoagulation [52]. In this method the negative charges on the GO sheets formed in mild alkaline solution contribute to the exfoliation of GO and to the nanocomposite formation driven by the emergence of electrostatic interaction between GO and TiO<sub>2</sub> colloidal particles, which possess a net positive surface charge density in aqueous electrolyte solutions below the pH of the point of zero charge ( $\text{pH}_{\text{p.z.c.}} = 6.3$ ) [53].

Our goal was to prepare GO/TiO<sub>2</sub> composite type materials from Brodie's GO by heterocoagulation in order to test them as catalysts in the methanol photocatalytic reforming reaction in liquid phase. To our knowledge, it is the first use of this type of composite in the

above reaction. Correlations between the photocatalytic behavior and the interaction of the carbonaceous material and the semiconductor were sought. Our recent results revealed that the working conditions of the methanol photocatalytic reforming reaction may result in significant changes of the structure of certain metal oxide-semiconductor catalyst systems involving both the semiconductor [54] and co-catalyst [15] compared to the fresh state. Therefore we characterized both the fresh and recovered samples by bulk and surface characterization methods such as diffuse reflectance UV-Vis,  $^{13}\text{C}$  MAS NMR, ATR-IR spectroscopy and XPS.

## 2. Experimental

### 2.1 Materials

Nanocomposites were prepared using commercially available P25  $\text{TiO}_2$  (Evonik, Germany). This solid contains 75% anatase and 25% rutile with a specific surface area ( $a_{\text{BET}}^{\text{S}}$ ) of  $50 \text{ m}^2/\text{g}$  and its average primary particle size is 30 nm (manufacturer data). GO was synthesized from natural flaky graphite (Graphitwerk Kropfmühl AG, Germany) by the Brodie method. This sample is highly oxidized ( $\text{C}_2\text{O}_{0.98}\text{H}_{0.40}$ ), and it is identical with that codenamed as GO-2 in an earlier publication [29]. Sodium hydroxide, hydrogen chloride used for pH setting and methanol, absolute ethanol solvent were products of Reanal (Hungary). Double distilled water (18 M $\Omega$ ) was used in every experiment.

### 2.2 Preparation of photocatalysts

A series of GO/ $\text{TiO}_2$  nanocomposite type photocatalysts with various GO content (1w%, 2w%, 10w%) were prepared from aqueous dispersions of P25  $\text{TiO}_2$  and exfoliated GO by heterocoagulation as described before [52]. Briefly, 40-400 mg of GO samples was dispersed in 1.6 L of water. In each case, the pH was adjusted to  $8.5 \pm 0.3$ . After 15 min of sonication and one day of continuous stirring in the dark the pH was reset to  $8.5 \pm 0.3$ . Next, these suspensions were poured into 0.4 L of 9–9.9 g/L aqueous colloid dispersions of  $\text{TiO}_2$  (pH=5) upon which the system coagulated in several seconds. However, we must note that the composites, unlike to the reported cases in ref [52], settled completely only after the first washing step and the supernatants remained slightly turbid even at larger GO concentrations when they were progressively washed from the electrolyte contaminations. The reason might have been the ageing of the GO samples, resulting in incomplete exfoliation. After decantation or centrifugation the wet sediment was dried at  $50^\circ\text{C}$  and then crushed into powder. Blank  $\text{TiO}_2$  was obtained by the same procedure in the absence of GO.

In order to simulate the effect of each assumed process which can contribute to the structural changes of the heterocoagulated GO/TiO<sub>2</sub> composites under the conditions of the photocatalytic reaction, samples for model experiments were prepared by certain treatments of the fresh GO/TiO<sub>2</sub> composite. In order to simulate the physical removal of GO sheets from the composites and their rupture, the fresh sample of 2w% GO/TiO<sub>2</sub> was stirred for a week at room temperature in the aqueous methanol reaction mixture in dark. This agitated sample was denominated as **A-GO/TiO<sub>2</sub>** and it was used in the same reaction mixture in which it was prepared. Removal of functional groups and reduction of carbon skeleton can also occur in the methanol solution during the photoinduced reaction. In order to get a model sample representing this process the fresh composite of 2w% GO/TiO<sub>2</sub> underwent a hydrothermal treatment at 120°C for 24h in an N<sub>2</sub> rinsed autoclave in ethanol/water 1:1 mixture similarly to the procedure used for composite preparation from Hummer's GO [46]. This hydrothermally treated sample was denominated as **HT-GO/TiO<sub>2</sub>**. Model sample **AHT-GO/TiO<sub>2</sub>** was obtained by agitation of HT-GO/TiO<sub>2</sub> and it also was used in the same reaction mixture in that it was stirred. In order to show the sole effect of UV irradiation during the photoinduced reaction the dry sample of the fresh 2w% GO/TiO<sub>2</sub> was exposed to UV irradiation (430 nm) for 2h. The UV irradiated model sample was denominated as **UV-GO/TiO<sub>2</sub>**.

### 2.3 Characterization of the photocatalysts

ATR-IR spectra were recorded by the means of a Varian 2000 (Scimitar Series) FT-IR spectrometer (Varian Inc, US) equipped with an MCT (Mercury-Cadmium-Telluride) detector and with a 'Golden Gate' diamond single reflection ATR unit (Specac Ltd, UK). 64 scans were collected at a spectral resolution of 4 cm<sup>-1</sup>. Baseline correction was performed using the GRAMS/AI (7.02) software (ThermoGalactic Inc., US).

Solid state NMR magic angle spinning (MAS) spectra of the fresh and recovered samples were recorded on a Varian NMR System (Varian Inc., Palo Alto, CA, U.S.A.) operating at <sup>1</sup>H frequency of 400 MHz with a Chemagnetics T3 4.0 mm narrow bore double resonance probe. The <sup>1</sup>H MAS spectra were measured with a rotor spinning rate of 12 kHz. The <sup>1</sup>H  $\pi/2$  pulse was 3  $\mu$ s and a repetition delay of 30 seconds was used. For the <sup>13</sup>C CP MAS (cross-polarization magic angle spinning) [55] spectra 20000 transients were recorded with 2 ms of contact time with SPINAL-64 decoupling [56] and 5 seconds of recycle delay. The direct polarization <sup>13</sup>C MAS spectra were recorded with 60 s of relaxation delay and 20000 transients were collected. For the <sup>13</sup>C spectra adamantane was used as external chemical shift reference (38.55 and 29.50 ppm),  $\pi/2$  pulse lengths were 3.0  $\mu$ s for carbon and 3.0  $\mu$ s for the

proton channel. For the  $^{13}\text{C}$  measurements a rotor spinning rate of 8 kHz were used. The measuring temperature was 20°C in all cases.

X-ray photoelectron spectroscopy (XPS) measurements were carried out using an EA125 electron spectrometer manufactured by OMICRON Nanotechnology GmbH (Germany). The photoelectrons were excited by  $\text{MgK}\alpha$  (1253.6 eV radiation. Spectra were recorded in the Constant Analyzer Energy mode of the energy analyzer with 30 eV pass energy resulting in a spectral resolution of around 1 eV. For XPS experiments the samples in the form of fine powder were suspended in isopropyl alcohol. Drops of this suspension were placed on standard OMICRON sample plates; after evaporation of the solvent catalyst coatings with sufficient adhesion and electric conductivity were obtained. Effects of possible electric charging were compensated by adjusting the binding energy of the Ti  $2p_{3/2}$  peak to 458.8 eV (consensual value for  $\text{TiO}_2$  [42]). Chemical states of the elements were deduced from high resolution spectra using XPS databases [57-58]. Quantification was performed using combination of CasaXPS [59] and XPS MultiQuant [60].

#### *2.4 Photocatalytic hydrogen generation*

The photocatalytic reaction was carried out in liquid phase in a reactor system of 10 quartz glass units equipped with magnetic stirrers, gas inputs and outputs as described before [15]. The size of the cylindrical glass units were: 60 mm in height and 140 mm in diameter. Nitrogen gas with 20 ml/min flow rate was continuously bubbled through all reactor units in parallel. Gas outlets were connected to the gas chromatograph (GC) via a ten position selector valve. According to blank experiments all of the reactor units were equivalent in terms of the catalytic activity. In case of kinetic measurements one channel mode was used. The initial concentration of methanol was 6 v % in distilled water. It has been known that the rate of hydrogen generation vs. methanol concentration relationship gives a saturation curve and use of diluted solution is favorable [16]. The reaction was carried out at room temperature. The amount of catalyst and the reaction volume in every unit was 0.100 g and 280 ml, respectively. Osram HQL de luxe 125W lamps were used as light sources operated in UV-visible region. The reaction was monitored for 4 h. Hydrogen formation was followed by GC analysis of the outlet gas upon using SUPELCO Carboxen 1010 column, TCD and FID detection and argon internal standard. The calculation of the  $\text{H}_2$  formation rate was based on the results obtained by TCD. FID was used to monitor the methanol signal, sudden decrease of which indicated if clogging from the water appeared in the gas tube system.

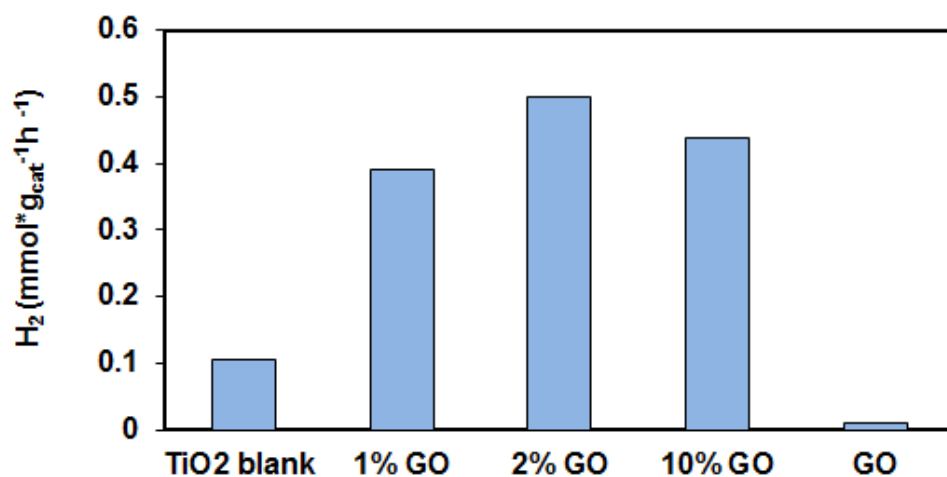


After the photocatalytic reaction, the samples were recovered from the aqueous methanol solution by centrifugation, washing with 3x50 ml absolute ethanol followed by drying under N<sub>2</sub> flow.

### 3. Results and discussion

#### 3.1 Behavior of the nanocomposites obtained by heterocoagulation

As products, only H<sub>2</sub> and CO<sub>2</sub> could be detected by TCD in the outlet gas. Although the liquid phase was not analyzed, it is known that methanol is photooxidized to CO<sub>2</sub> via the formation of the stable intermediates [4]. Consequently, presence of formaldehyde, formic acid, methyl formate, and CO<sub>2</sub> can be supposed in the diluted aqueous solution besides the starting MeOH. Regarding the photocatalytic behavior in the methanol reforming reaction, the introduction of GO into the composite increased the catalytic activity about three-to-fourfold compared to the blank TiO<sub>2</sub> while unsupported GO led to H<sub>2</sub> formation at the detection limit (see Fig. 1). This increase is slightly less than that reported by El-Bery et al. [61]. The H<sub>2</sub> evolution showed a maximum as a function of the GO content, similarly as reported before [42, 49]. For comparison, significantly smaller hydrogen formation rates were observed with 1.0wt% Ag/TiO<sub>2</sub> reference catalysts prepared by sodium borohydride reduction, while a 0.5wt% Pt/TiO<sub>2</sub> reference catalyst obtained by impregnation and subsequent reduction in H<sub>2</sub> gave outstandingly better results (see data of H<sub>2</sub> formation and details of preparation of reference catalysts in the Electronic Supplementary Materials). We must note here also that, for the fair comparison of activities expressed in mmol g<sup>-1</sup> h<sup>-1</sup> dimension, the specific surface area of the catalysts needs to be similar to each other. Here, this criterion is fulfilled because the composite particles are characterized by a large mass excess of TiO<sub>2</sub>. Although graphite oxide particles have a large specific surface area at high degrees of exfoliation, the surface area of the composites must fall close to that of the TiO<sub>2</sub> particles ( $a_{\text{BET}}^{\text{S}} = 50 \text{ m}^2/\text{g}$ ) that completely cover the surface of the delaminated carbon sheets at low GO loadings, as shown in our previous publication [52]. Moreover, although the total surface area may be slightly higher than for the bare particles, the reactive surface area is only represented by the TiO<sub>2</sub> particles, which is exactly the same for both the “TiO<sub>2</sub> blank” and for the nanocomposite samples irrespectively of their phase ratios.



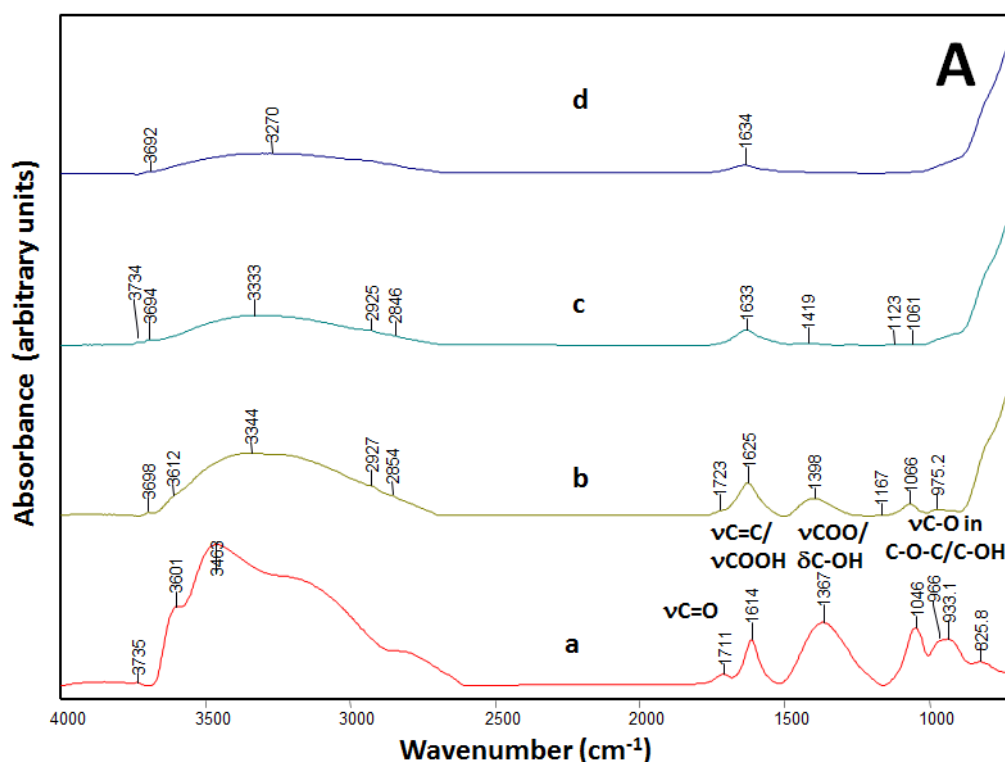
**Fig. 1** Rate of H<sub>2</sub> formation after 240 min reaction time

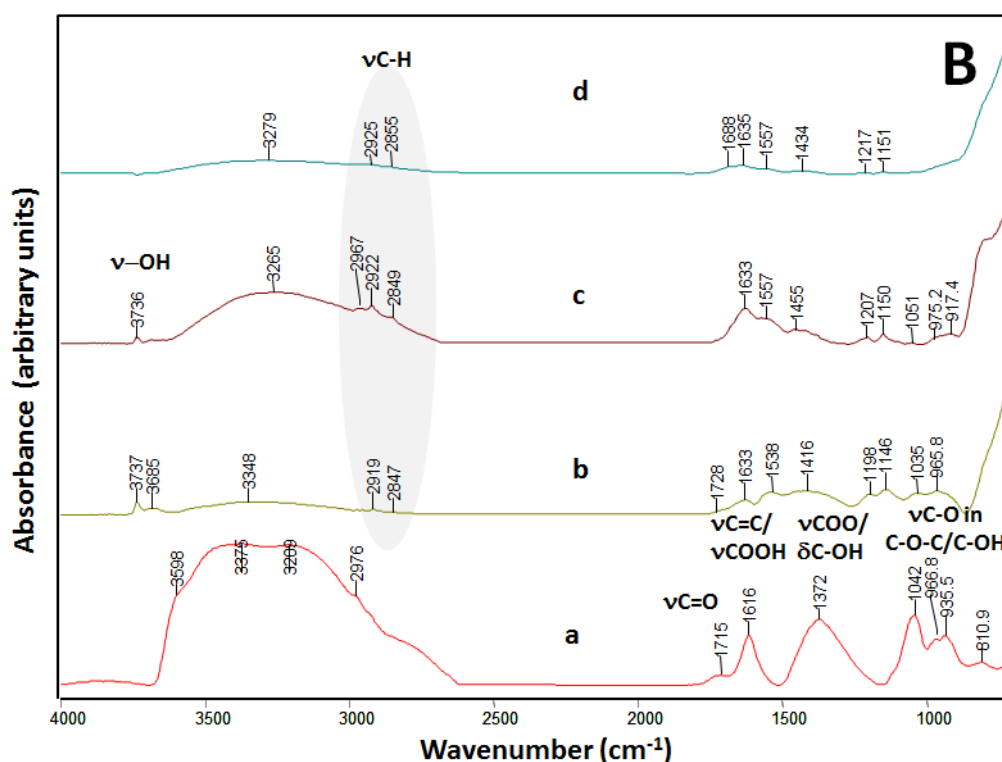
During the photocatalytic reaction significant color change (blackening) of the GO/TiO<sub>2</sub> composites was observed. The UV-visible spectroscopic measurements of fresh and recovered composite samples confirmed the darkening (see Fig. 1SA in the Electronic Supplementary Materials). The absorbance of the fresh GO/TiO<sub>2</sub> samples was proportional to the carbon content (see Fig. 1SA in the Electronic Supplementary Materials) as has been described in the literature [26, 49]. At the same time the color change of pristine TiO<sub>2</sub> and blank TiO<sub>2</sub> (see diffuse reflectance UV-Vis spectra in Fig. 1SB in the Electronic Supplementary Materials) during the irradiation in methanol solution was negligible. The slight grey color can be explained by the organic deposits picked up from the solution. According to photoelectron spectroscopy measurements, no reduction of the TiO<sub>2</sub> particles is evident after the photocatalytic reaction, thus the significant blackening of the recovered composite type samples indicates that structural changes of the GO/TiO<sub>2</sub> composites occurred during the reaction and these changes must be related to the GO part.

In order to map the supposed structural/chemical changes of GO, ATR-IR spectra of fresh (Fig. 2A) and recovered samples (Fig. 2B) were recorded. It has to be noted that unsupported GO and 10w% GO/TiO<sub>2</sub> had rather dark color, resulting in a heavy baseline shift so all the discussed spectra underwent baseline correction. It can be safely reported that composite preparation by heterocoagulation method resulted in only little changes in the structure of GO, as all major spectral features are observable after the procedure (*cf.* lines a and b in Fig. 2A). In details, pure GO exhibits several bands corresponding to oxygen containing surface species like –OH groups ( $\nu$ -OH stretchings around 3735, 3600 and 3460 cm<sup>-1</sup>), carbonyl group ( $\nu$ C=O at 1711 cm<sup>-1</sup>) and carboxylates ( $\nu$ COO at 1641 and 1367 cm<sup>-1</sup>).

The latter may be overlapped with C=C stretching of aromatic rings. The bands between 1000 and 800  $\text{cm}^{-1}$  belong to C-O stretching vibrations of C-OH and C-O-C species. The spectrum resembles that of pure  $\text{TiO}_2$ ; only very small bands referring to oxygen-containing surface species can be witnessed (surface -OH and surface carbonates/carboxylates at 3734  $\text{cm}^{-1}$  and at 1419, 1123 and 1061  $\text{cm}^{-1}$ , respectively).

As far as the recovered samples are concerned, little changes are seen in the case of the pure GO. On the other hand, in the case of the used  $\text{GO}/\text{TiO}_2$  catalysts, several bands indicate species derived from adsorbed MeOH/organic compounds on  $\text{TiO}_2$  (Fig. 2B, lines b, c and d) which resulted in difficulty in the interpretation of the spectra.





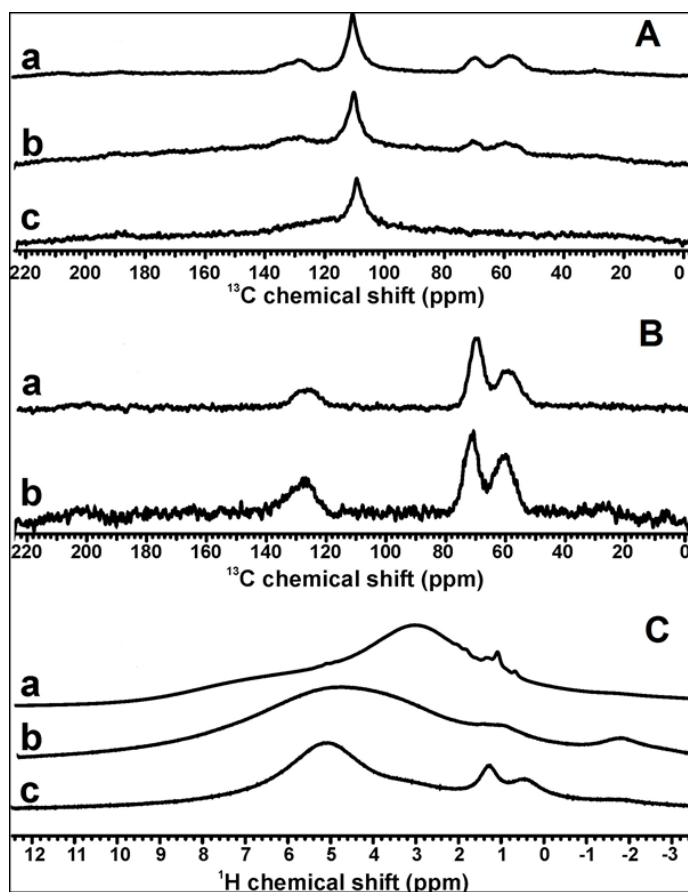
**Fig. 2** ATR-IR spectra of fresh (A) and recovered samples (B)

a: GO; b: 10w% GO/TiO<sub>2</sub>; c: 2w% GO/TiO<sub>2</sub>; d: TiO<sub>2</sub>

Accordingly, ATR-IR analysis of the recovered samples was not really informative for the changes of the supported GO but it is important to mention that newly formed OH groups (at 3736 and 3685 cm<sup>-1</sup>) in the GO containing recovered samples were well-marked (see lines b and c in Fig. 2B). It seems reasonable that the GO activates the TiO<sub>2</sub> surface.

In order to get further information exclusively about the carbonaceous part of the fresh and recovered samples, <sup>13</sup>C solid state MAS NMR technique was chosen. Because of the low carbon content of the samples only the composite sample of 10w% GO/TiO<sub>2</sub> and unsupported GO were investigated in detail. Direct polarization spectra of fresh GO/TiO<sub>2</sub> composite and bulk GO are very similar to each other (*cf.* spectrum a and b in Fig. 3A) confirming the finding with ATR-IR, *i.e.* no significant structure change of GO appeared during the composite preparation by heterocoagulation. Characteristic functional groups on the carbon skeleton such as primary alcohol (55-65 ppm), secondary alcohol and ether (65-75 ppm), and oxo (195-215 ppm) groups could be identified. Beside them signal of isolated C=C and polyaromatic part (120-140 ppm) could be observed. The strong signal at 110 ppm belongs to the Teflon parts of probe. Cross polarization MAS NMR technique enhances the intensity of

carbon signals and removes the Teflon background. This enhancement is weaker for quaternary carbons and it is suitable for direct quantitative analysis. Signals in the region of 50-80 ppm belonged to CH connected with heteroatom (O) and the edge of the graphite sheets as C-H signals in these spectra were more intense than that of quaternary or condensed aromatic carbons.



**Fig. 3** MAS NMR spectra of GO and GO/TiO<sub>2</sub> composite samples A: <sup>13</sup>C MAS NMR spectra; B: <sup>13</sup>C MAS NMR cross polarization spectra; C: <sup>1</sup>H MAS NMR  
a: GO; b: fresh 10w% GO/TiO<sub>2</sub>; c: recovered 10w% GO/TiO<sub>2</sub>.

Contrary to the fresh composite, the recovered sample showed significant differences compared to starting GO; it lost its functional groups (*cf.* spectrum a and c in Fig. 3A). Furthermore, significant decrease of the carbon content was indicated by the less intense signals and the need of the increased time of measurement. In addition, CP spectra could not be recorded over this sample which suggested that magnetic susceptibility and/or dielectric parameters of the 10w% GO/TiO<sub>2</sub> composite also changed drastically.

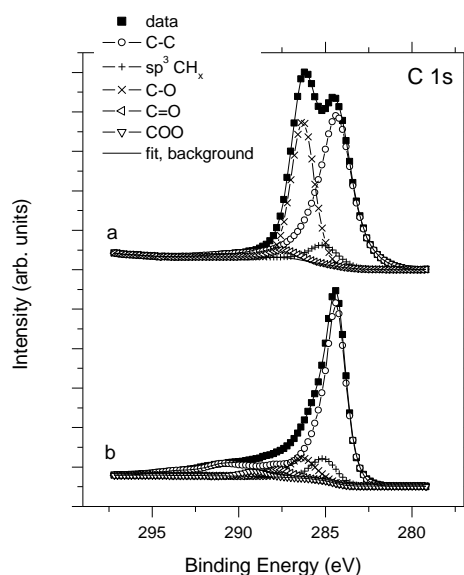
Results of <sup>1</sup>H MAS NMR measurements are presented in Fig. 3C. Signals of aromatic (~7 ppm) and to heteroatom connected aliphatic (~3.5 ppm) protons can identified

on the spectrum of GO (line a in Fig. 3C). Spectra of fresh and recovered composites (lines b, c in Fig. 3C) fit to the characteristic spectrum of  $\text{TiO}_2$  [62]. In the composites water could be present related to  $\text{TiO}_2$  by adsorption and to GO sheets by intercalation which cannot be distinguished. In case of the recovered sample the signal intensity decreases and at 5 ppm in the region of bridged OH groups and adsorbed water a narrower band appears, while the loosely bonded broad water signal vanishes. This observation suggests that the recovered composite sample has more hydrophobic character than the fresh one in accordance with the removal of O-containing functional groups which was indicated by  $^{13}\text{C}$  MAS NMR. Parallel, newly formed terminal OH groups (0-2 ppm) can provide certain hydrophilicity for the composite.

Composition of the parent GO and the fresh and recovered photocatalysts as well as the bonding environment of the components were also investigated by X-ray photoelectron spectroscopy. Fig. 4 shows the C 1s XPS spectra of unsupported GO in the as received state (line a) and after reduction in  $\text{H}_2$  for 1h (line b). Deconvolution of the spectra was carried out by considering quantitative relationships with the corresponding O 1s spectra [63, 64]. In the case of the initial GO sample, the two-peaked spectrum contains a broad, slightly asymmetric feature around 284.4 eV, which can be assigned to carbon atoms participating in C-C bonds. Its large width probably indicates a distribution of slightly different local environments due to the disturbing effect of the neighboring oxygen-containing functional groups [64]. The other main feature around 286.2 eV is due to carbon atoms singly bonded to oxygen, such as in  $\text{sp}^2$  or  $\text{sp}^3$  C-OH or C-O-C in cyclic ethers. A small peak around 285.0 eV can be assigned to  $\text{sp}^3$   $\text{CH}_x$  species or defect sites in aromatic environments [64]. Finally, a weak contribution around 287.4 eV indicates the presence of carbonyl functionalities. XPS data are in good agreement with NMR as both techniques indicate alcoholic and carbonyl-like functional groups as the main oxygen-containing moieties in GO.

Reduction at 300 °C in  $\text{H}_2$  for 1 h results in a very significant loss of oxygen-containing functional groups of unsupported GO and leads to the appearance of a graphite-like spectrum. This can be explained well by the complete thermal decomposition of graphite oxide because the annealing temperature exceeds the deflagration temperature of Brodie-GO [29] of ca. 275 °C. The asymmetric main component of the spectrum of Fig. 4 arises at 284.4 eV and is accompanied by characteristic energy loss features shifted by some 6 eV from the main peak as a result of  $\pi$ - $\pi^*$  excitations in the aromatic system. Nevertheless, small contributions around 285.0 eV (hydrocarbon-like moieties or defect sites in the graphite planes), 286.3 eV (C-O bonds in alcoholic or ether groups), 287.5 eV (C=O groups) and 289.1

eV (carboxylic groups) are still present. These structural parameters are in accordance with the common observation that the carbon material formed upon thermal degradation still contains oxygenated functional groups [64, 65].

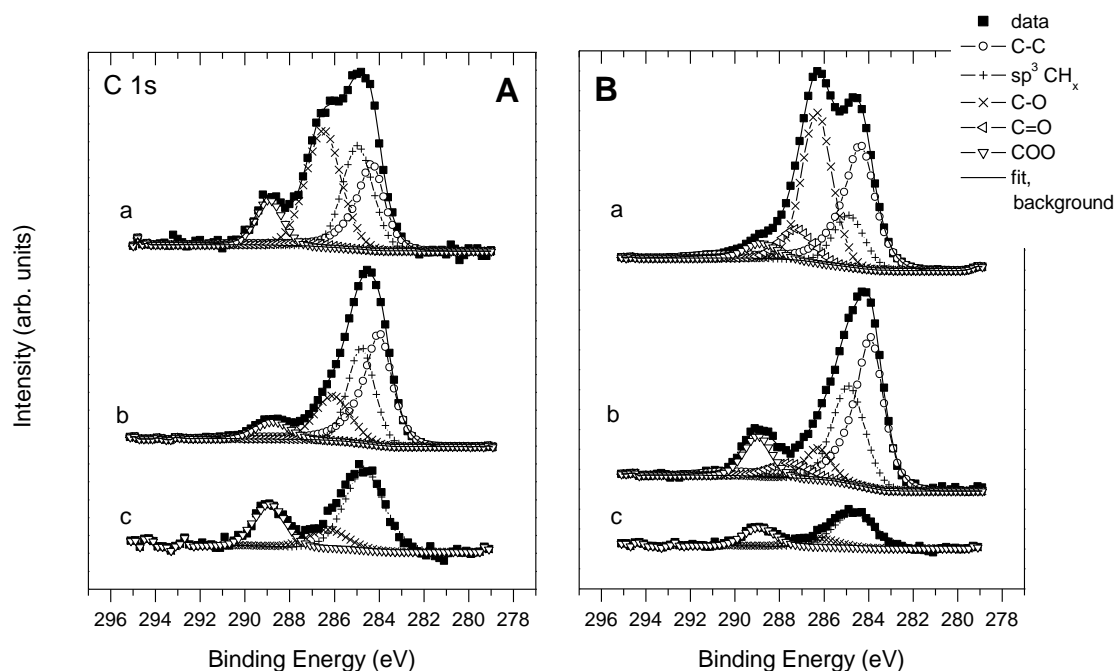


**Fig. 4** C 1s XPS spectra of GO in its initial state (a) and after reduction at 300 °C in hydrogen for 1 h (b)

In Fig. 5 C 1s spectra of the 2% (A) and the 10w% GO/TiO<sub>2</sub> (B) samples are compared before and after the photocatalytic process. Evaluation of the data is complicated by the fact that even the pristine P25 powder contains a significant amount of carbonaceous contaminants. The C 1s spectrum of this contamination is shown after scaling according to the C/Ti ratio as the lowermost curves both in (A) and (B) panels. A quantitative evaluation of the composition of the samples is given in Table 2.

Table 2. A quantitative evaluation of the composition of the samples based on XPS data

Sample	state	Composition (atomic %)		
		C	Ti	O
2w% GO/TiO <sub>2</sub>	as synthesized	14	25	61
	recovered	24	23	53
10w% GO/TiO <sub>2</sub>	as synthesized	32	17	51
	recovered	12	26	62
P25	as received	8	28	63



**Fig. 5** C 1s XPS spectra of the 2w% (A) and the 10w% GO/TiO<sub>2</sub> (B) samples in the fresh (curve a) and recovered (curve b) state. C 1s region of the bare P25 TiO<sub>2</sub> is also shown indicating the presence of adventitious hydrocarbon contamination (curve c, intensity is scaled according to the C/Ti ratio).

Spectra of the GO/TiO<sub>2</sub> samples were analyzed by deconvolution using the components found already in the spectra of the original and the reduced GO. This analysis suggests that the main contribution to the spectra of the fresh samples in both the 2w% and 10w% cases comes from a double-peaked feature with maxima at 284.3 and 286.3 eV, which corresponds to the C-C and C-O peaks from the spectrum of GO. The relatively strong peak at 284.9 eV and the weaker one at 288.9 eV in the spectrum of the 2w% sample can be assigned to the contamination on the TiO<sub>2</sub> (hydrocarbon and carboxyl, respectively). The same contamination-related peaks appear at a much smaller relative intensity in the case of the as synthesized 10w% sample, in agreement with its expected higher GO content. The fact that GO-like features can be identified in the C 1s spectra of the GO/TiO<sub>2</sub> photocatalysts confirms that the exfoliation resulted in no drastic change of the functionalities of GO, as already indicated by ATR-IR and NMR.

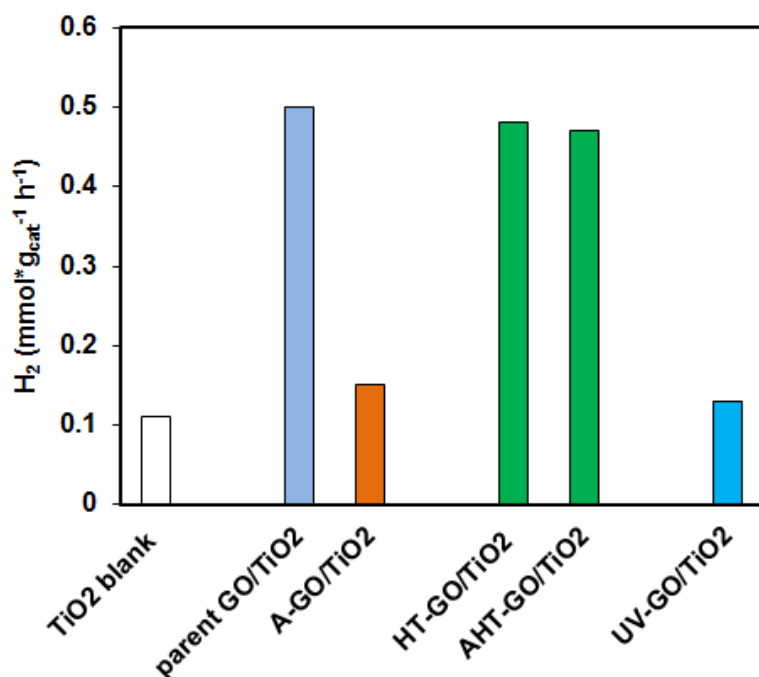
Regardless to the initial GO content, very similar changes are observed in the recovered samples. A very significant loss of the C-O-type contribution (286.3 eV) occurs and



the graphite-like C-C contribution becomes the dominating spectral feature along with a shift to binding energies slightly below 284.0 eV. The hydrocarbon peak at 284.8 eV remains strong. The observed changes are qualitatively similar to those reported previously for GO/TiO<sub>2</sub> photocatalysts used for photocatalytic degradation of phenol [52]. In case of the 2w% sample the carbon content somewhat increased during the photocatalytic reaction, while a drastic decrease was found in the 10w% case.

### 3.2 Model experiments for identifying the processes leading to structural transformation of the GO/TiO<sub>2</sub> composite during photocatalytic methanol reforming

Comparison of fresh and recovered samples by UV-Vis, MAS NMR and XPS techniques (Chapter 3.1) revealed that the GO/TiO<sub>2</sub> composite did not keep its structure under the conditions of the photocatalytic reaction. Physical removal of GO sheets during the long time stirring was just as imaginable as the reduction of GO by the alcohol in the slurry and/or by the reaction product *i.e.* the *in situ* formed hydrogen. UV-induced degradation of the carbonaceous material also could not be excluded.



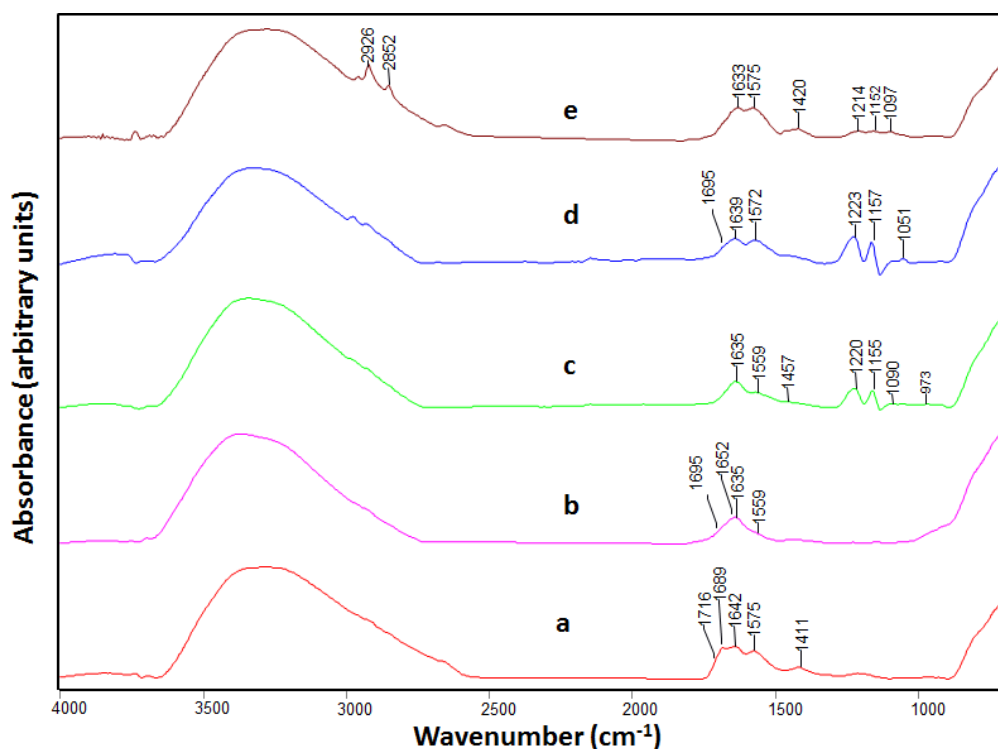
**Fig. 6** Rate of H<sub>2</sub> formation in model experiments with specially treated catalysts after 240 min reaction time. For abbreviations see the experimental part

In an attempt to study the effect of these probably parallel processes separately, catalyst samples were specially treated to model their influence (see experimental part). Photocatalytic behavior of the model system prepared by agitation in the non-illuminated

reaction mixture (A-GO/TiO<sub>2</sub> (Fig. 6)) indicated that long term stirring of the slurry in dark had significant negative effect on the activity of H<sub>2</sub> formation in accordance with the supposed mechanical disintegration of the composite. The sample obtained using UV-treated GO (UV-GO/TiO<sub>2</sub>) showed also strongly decreased activity, while the hydrothermal treated sample (HT-GO/TiO<sub>2</sub>) was as active as the starting composite prepared by heterocoagulation. During the hydrothermal treatment the composite obtained by heterocoagulation was transformed into a stable form, because it maintained the activity of the parent composite; even long-term agitation did not reduce its activity (see AHT-GO/TiO<sub>2</sub> in Fig. 6). These findings strongly imply that the strength of electrostatic forces between the oppositely charged TiO<sub>2</sub> particles and GO platelets decrease in aqueous methanol solution in contrast to pure water phase. The presence of the alcohol changes both the dielectric and solvation properties of the medium and causes a significant effect on the structure of the electric double layer developed at the surface of charged colloidal particles. However, the fact the hydrothermally treated samples retain their nanostructure shows that under these conditions stronger bonds form between the counterparts of the composite sample.

Regarding the color of fresh and recovered model samples, the most striking observation was the appearance of dark color in the fresh HT-GO/TiO<sub>2</sub> which then did not change significantly during the photocatalytic reaction. The other model samples behaved very similarly to the parent 2w% GO/TiO<sub>2</sub>; strong darkening was observed during the photocatalytic reaction regardless of the degree of development of the H<sub>2</sub>. For further details see the Electronic Supplementary Materials.

ATR-IR spectra of model catalysts are presented in Fig. 7.



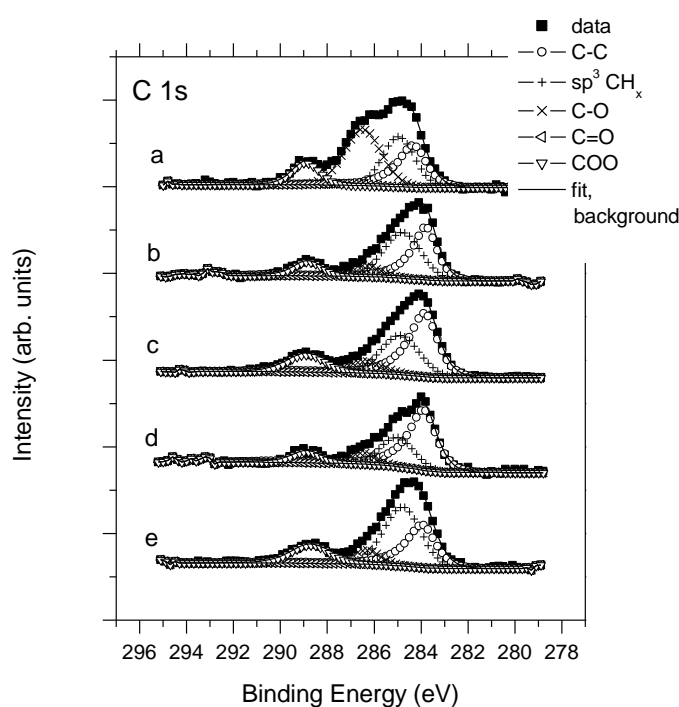
**Fig. 7** ATR-IR spectra of model catalysts

a: fresh UV-GO/TiO<sub>2</sub>, b: fresh HT-GO/TiO<sub>2</sub>, c: recovered A-GO/TiO<sub>2</sub>, d: recovered AHT-GO/TiO<sub>2</sub>, e: recovered UV-GO/TiO<sub>2</sub>. For abbreviations see the experimental part

It can be seen that UV irradiation of the parent composite in dry condition (*cf.* line b in Fig. 7 and line c in Fig. 3A) removed certain functional groups of the initial sample (-OH, C=O, C-O). Agitation did not result in significant change of the spectra of the recovered composite (*cf.* line c in Fig. 7 and line b in Fig. 3B), however, the relative intensity of  $\nu$ C-O bands (1250-800 cm<sup>-1</sup>) was increased (for comparison all spectra were normalized to the highest Ti-O band around 628 cm<sup>-1</sup>; not shown). It seems plausible that the GO framework fractures were followed by formation of end C-OH groups. The same phenomenon was witnessed also in the case of hydrothermal treated agitated and recovered sample (see line d in Fig. 7). Moreover, the relative intensity of  $\nu_{as}$ COO<sup>-</sup> (around 1572 cm<sup>-1</sup>) of recovered UV and HT treated spectra is also increased, suggesting a small change in GO geometry.

In Fig. 8 C 1s XPS spectra of the model samples recovered after the photocatalytic methanol reforming reaction are compared with that of the parent 2w% GO/TiO<sub>2</sub> catalyst. In addition, spectrum for the fresh HT-GO/TiO<sub>2</sub> is also shown. The spectra of the recovered catalysts are generally similar, containing mainly graphitic and hydrocarbon-like contributions; the relative amount of oxidized carbon species becomes small. This overall

reduction correlates well with the observed darkening of all model samples. Even agitation in the reaction mixture is enough for disruption of the original GO structure, which results in a very significant loss of the C-O functionalities. In the spectra of the hydrothermally treated catalysts both in the fresh and the recovered state the dominant contribution arises from graphitic carbon; its binding energy around 284.0 eV suggests the pronounced coupling of the carbon skeleton to TiO<sub>2</sub> during the treatment. According to the XPS data, the least graphitic system is the UV-illuminated sample; the spectra suggest that illumination in the absence of the reaction mixture produces hydrocarbon-like species, which have lost their cocatalytic activity. The carboxyl-like contribution around 289 eV can most probably be assigned to the inevitable adventitious hydrocarbon contamination.



**Fig. 8** XPS spectra of the 2w% GO/TiO<sub>2</sub> samples used in the model experiments

a: parent 2w% GO/TiO<sub>2</sub>; b: recovered A-GO/TiO<sub>2</sub>; c: fresh HT-GO/TiO<sub>2</sub>; d: recovered HT-GO/TiO<sub>2</sub>; e: recovered UV-GO/TiO<sub>2</sub>. For abbreviations see the experimental part

### 3.3 Role of GO/TiO<sub>2</sub> transformation during the photocatalytic reaction

Comparison of fresh and recovered composites obtained by heterocoagulation (Chapter 3.1) showed that the liquid phase working form of the catalyst evolved *in situ* under

the conditions of the photocatalytic reaction. The most important changes affected the GO part of the composite. These were the removal of the O containing functional groups (demonstrated by  $^{13}\text{C}$  MAS NMR results) and graphitization (shown by XPS). It's worth mentioning that GO lost its functional groups by UV irradiation in model catalyst UV-GO/TiO<sub>2</sub> similarly to GO/TiO<sub>2</sub> did under the photocatalytical conditions, however the beneficial effect in the former case was missing. Most likely the remaining carbon skeleton was different.

It also could be concluded that transformation of GO itself was not the only factor in the achievement of higher H<sub>2</sub> formation compared to the bare TiO<sub>2</sub>. As the darkening in the recovered model samples did not bring increased H<sub>2</sub> production in every case, a specific transformation had to be assumed.

Results of  $^{13}\text{C}$  MAS NMR and XPS measurements on recovered samples presented loss of carbon content indicating that certain physical degradation occurred due to agitation under the photocatalytical conditions of (*i.e.* long time stirring in aqueous reaction mixture) which was confirmed by the photocatalytic behavior of model sample A-GO/TiO<sub>2</sub>. Despite the above effect which had obviously negative influence on the H<sub>2</sub> production, the composite sample obtained by heterocoagulation kept its activity. This probably happened because other factors had offset; such a factor could be the formation of a strong Ti-O-C interaction. According to the literature this kind of strong Ti-O-C interaction develops on points of the GO sheets from where functional groups were removed by the hydrothermal treatment [46-47]. In addition, the UV-assisted photocatalytic reduction of graphene oxide for preparation of R-GO/TiO<sub>2</sub> composites is a well-known method [66], during which TiO<sub>2</sub> and GO is irradiated in a slurry in a reductive medium *i.e.* ethanol. We believe in the UV-treated model sample (UV-GO/TiO<sub>2</sub>), depletion of functional groups did not allow such a linkage in the subsequent reaction and/or the damaged carbon skeleton prevented the formation of the preferred RGO structure. Similarly, agitation in the non-illuminated reductive reaction mixture may have resulted in too much damage in the carbon structure, so the beneficial Ti-O-C coupling only partly formed during the subsequent photocatalytic process.

In the hydrothermally treated composite sample (HT-GO/TiO<sub>2</sub>) the structure coupled with electrostatic interactions was transformed into TiO<sub>2</sub>-bound RGO prior to the photocatalytic reaction. This strong connection provided stability against mechanical effects too, which explained that model sample AHT-GO/TiO<sub>2</sub> was as active as the starting one. Moreover, the differences in the hydrogen formation rate - reaction time dependencies which were obtained upon using the parent 2%GO/TiO<sub>2</sub> or the hydrothermal treated (HT-GO/TiO<sub>2</sub>)

photocatalysts (see chapter “2.4*S Time dependence of H<sub>2</sub> production rate*” in the Electronic Supplementary Materials) provided evidences for the *in situ* changes of the GO/TiO<sub>2</sub> catalysts during the photocatalytic reaction.

ATR-IR results suggest that presence of GO contributes to the activation of TiO<sub>2</sub>; ordered OH groups of TiO<sub>2</sub> appeared in the recovered samples prepared by heterocoagulation. Cao and coworkers [67] reported that OH-TiO<sub>2</sub>@TiO<sub>2</sub> nanocrystal heterojunctions formed on RGO sheets enhanced the photocatalytic activity. In another work hydroxylation of the semiconductor was mentioned to prevent the closure of the photocatalytic cycle [68]. Further investigations are needed to decide whether the newly formed ordered OH groups in the recovered heterocoagulated samples, which were absent in the bare TiO<sub>2</sub> and the catalysts treated for the model experiments, played a certain role in the activity enhancement or not.

Although RGO/TiO<sub>2</sub> composites have been widely used as photocatalysts for different reactions, a new study proved that RGO was not a universal promoter [69]. The factor of photocatalytic activity enhancement was strongly affected by the various functional groups appearing in the organic substances. The length of alkyl chain in alcohols and carboxylic acids had minimal influence on the overall activity while the number of hydroxyl groups could promote the activity enhancement in the presence of RGO. The enhancement factor for RGO/TiO<sub>2</sub> (5w% loading) in the mineralization of MeOH was only 1.1 [69]. The enhancement effect of the carbonaceous part in our composites was also relatively moderate compared to that of Pt co-catalyst (near 20 times less). However introduction of a small amount of Pt could result in a synergistic effect [26], this way the Pt demand could be reduced significantly. Consequently, the GO may become an important material which could make the photocatalyst cheaper; that's why understanding its behavior in composites is essential.

#### 4. Summary

GO/TiO<sub>2</sub> composites prepared by heterocoagulation from Brodie type GO were active as catalyst in the photocatalytic reforming of methanol. Diffuse reflectance UV-Vis, ATR-IR, <sup>13</sup>C MAS NMR and XPS spectroscopic measurement of fresh and recovered samples revealed that these composites underwent structural transformation (such as mechanical degradation, reduction and graphitization) during the photo-induced reaction. Model experiments with specially pre-treated photocatalysts suggested that *in situ* development of a strong chemical coupling between the reduced graphene oxide sheets and TiO<sub>2</sub> may be the key factor for the beneficial co-catalytic effect of the GO. This process is analogous to the strong Ti-O-C formation reported in RGO/TiO<sub>2</sub> composites obtained by UV-assisted photocatalytic

reduction in ethanol, which is confirmed by the fact that the fresh GO/TiO<sub>2</sub> composites could be transformed into highly active and more stable form by hydrothermal treatment.

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# SUPPLEMENTARY MATERIAL

## Graphite oxide-TiO<sub>2</sub> nanocomposite type photocatalyst for methanol photocatalytic reforming reaction

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### 1S. Experimental

#### 1.1S Preparation of reference catalysts

Pt(NH<sub>3</sub>)<sub>4</sub>(NO<sub>3</sub>)<sub>2</sub> (Aldrich) was used as a precursor of supported Pt nanoparticles. AgNO<sub>3</sub> was purchased from Molar (Hungary) as analytical grade chemical.

In order to prepare reference Pt/TiO<sub>2</sub> catalyst P25 was impregnated by aqueous solution of Pt(NH<sub>3</sub>)<sub>4</sub>(NO<sub>3</sub>)<sub>2</sub>. The dried sample with 0.5% w Pt load was reduced for 1 hour at 400°C in H<sub>2</sub> flow. The Ag/TiO<sub>2</sub> reference catalysts were made by mixing the dispersion of TiO<sub>2</sub> and aqueous solution of AgNO<sub>3</sub>. Silver was reduced by adding freshly made sodium borohydride solution (5:1 molar ratio (4x excess) in respect of silver) to the catalyst dispersions. After centrifuging the samples were dried at 50°C then crushed into powder.

#### 1.2S Sample preparation for model experiments

In order to simulate the effect of each assumed process which can contribute to the structural changes of the heterocoagulated GO/TiO<sub>2</sub> composites under the conditions of the photocatalytic reaction, model samples were prepared by certain treatments of the fresh GO/TiO<sub>2</sub> composite. In order to simulate the physical removal of GO sheets from the composites and their rupture, the fresh sample of 2w% GO/TiO<sub>2</sub> was stirred for a week at room temperature in the aqueous methanol reaction mixture in dark. This agitated sample was denominated as **A-GO/TiO<sub>2</sub>** and it was used in the same reaction mixture in which it was prepared. Removal of functional groups and reduction of carbon skeleton can also occur in the methanol solution during the photoinduced reaction. In order to get a model sample representing this process the fresh composite of 2w% GO/TiO<sub>2</sub> underwent a hydrothermal

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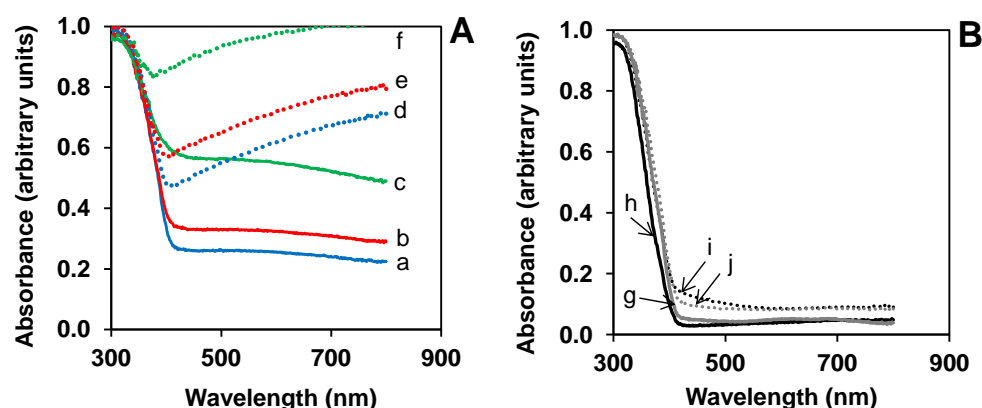
treatment at 120°C for 24h in an N<sub>2</sub> rinsed autoclave in ethanol/water 1:1 mixture similarly to the procedure used for composite preparation from Hummer's GO [1]. This hydrothermally treated sample was denominated as **HT-GO/TiO<sub>2</sub>**. Model sample **AHT-GO/TiO<sub>2</sub>** was obtained by agitation of HT-GO/TiO<sub>2</sub> and it also was used in the same reaction mixture in that it was stirred. In order to show the sole effect of UV irradiation during the photoinduced reaction the dry sample of the fresh 2w% GO/TiO<sub>2</sub> was exposed to UV irradiation (430 nm) for 2h. The UV irradiated model sample was denominated as **UV-GO/TiO<sub>2</sub>**.

### 1.3S Sample characterization

Diffuse reflectance UV-visible spectra of the samples were registered using a Jasco V-570 UV-VIS spectrophotometer equipped with a NV-470 type integrating sphere. The data were collected between 800 and 200 nm wavelengths with 100 nm/min speed.

## 2S. Results

### 2.1S Characterization of the fresh and recovered photocatalysts



**Fig. 1S** Diffuse reflectance UV -Vis spectra of the fresh and recovered samples

A: GO/TiO<sub>2</sub> composites; B: TiO<sub>2</sub>;

a: fresh 1w% GO/TiO<sub>2</sub>; b: fresh 2w% GO/TiO<sub>2</sub>; c: fresh 10w% GO/TiO<sub>2</sub>; d: recovered 1w% GO/TiO<sub>2</sub>; e: recovered 2w% GO/TiO<sub>2</sub>; f: recovered 10w% GO/TiO<sub>2</sub>; g: pristine TiO<sub>2</sub>; h: blank TiO<sub>2</sub> fresh; i: recovered TiO<sub>2</sub>; j: recovered blank TiO<sub>2</sub>

Fig. 1S shows the diffuse-reflectance UV-Vis spectra of the fresh and recovered photocatalysts. The absorbance of the fresh GO/TiO<sub>2</sub> samples was proportional to the carbon content (*cf.* line a, b, c in Fig.1S) as has been described in the literature [2, 3]. The recovered samples showed much higher absorption than the fresh ones, especially in the visible region and these values were also proportional to the initial carbon content of the composites (*cf.* line d, e, f in Fig. 1S). At the same time the color change of pristine TiO<sub>2</sub> and blank TiO<sub>2</sub> (see Fig.

S2B) during the irradiation in methanol solution was negligible. The slight deviation in the absorbance can be explained by the organic deposits picked up from the solution.

### 2.2S $H_2$ production over reference catalysts

Comparing the activity of the reference catalysts which contained metal co-catalyst (Table 1S) outstandingly high  $H_2$  evolution rate value was obtained in the presence of Pt as expected. However the hydrogen production of GO/TiO<sub>2</sub> ( $0.50 \text{ mmol} \cdot \text{h}^{-1} \text{g}_{\text{cat}}^{-1}$ ) exceeded that of Ag/TiO<sub>2</sub>.

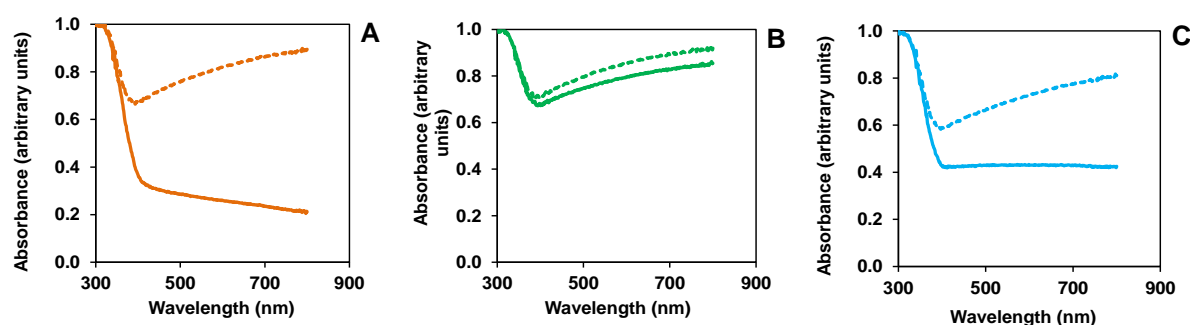
**Table 1S.** Comparison of  $H_2$  evolution rate over various P25 TiO<sub>2</sub> based photocatalysts

No	Co-catalyst	Load, w%	$H_2$ , $\text{mmol} \cdot \text{h}^{-1} \text{g}_{\text{cat}}^{-1}$
1 <sup>a</sup>	no (pristine)	no	0.12
2	Pt	0.5	9.29
3	Ag	1.0	0.30
4 <sup>b</sup>	Ag	1.0	0.29

<sup>a</sup>average of 3 parallel experiments

<sup>b</sup>parallel preparations

### 2.3S Characterization of the samples for model experiments



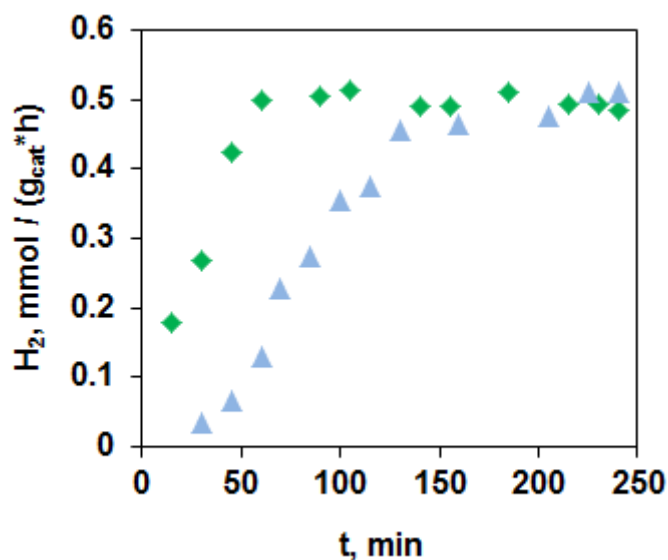
**Fig. 2S** Diffuse reflectance UV-Vis spectra of the fresh and recovered samples for model experiments

A: A-GO/TiO<sub>2</sub>, B: HT-GO/TiO<sub>2</sub>, C: UV-GO/TiO<sub>2</sub> (Abbreviations see in the experimental part). Solid line: fresh composite; dashed line: recovered sample

Regarding the UV-vis spectra of fresh and recovered model catalysts (Figure 2S) the most striking observation was the dark color of the fresh HT-GO/TiO<sub>2</sub> which then did not change significantly during the photocatalytic reaction. The other model samples behaved very similarly to the parent 2w% GO/TiO<sub>2</sub>; strong darkening was observed during the photocatalytic reaction regardless of the degree of development of the  $H_2$ .

### 2.4S Time dependence of $H_2$ production rate

The time dependence of the hydrogen formation rate of 2% GO/TiO<sub>2</sub> (parent composite) and HT-GO/TiO<sub>2</sub> (hydrothermal treated composite, see in Chapter 1.2S) is compared in Figure 3S:. Hydrogen formation rate of these samples after 240 min reaction time is given in Figure 6 of the article.



**Figure 3S.** Hydrogen formation rate - reaction time dependencies.

▲ - 2% GO/TiO<sub>2</sub> (parent composite)

◆ - HT-GO/TiO<sub>2</sub> (hydrothermal treated composite).

In our experiments the  $H_2$  formation rate - reaction time dependencies had two main parts, the initial period and a plateau. The plateau was reached earlier in case of HT-GO/TiO<sub>2</sub> (hydrothermally pretreated sample) than in case of the parent photocatalyst. In addition the slope is steeper in case of HT-GO/TiO<sub>2</sub> than in case of the parent photocatalyst. In our experimental system the plateau of the  $H_2$  evolution rate represented a continuous formation of  $H_2$  at a constant rate. The gradual increase in the hydrogen production rate during the initial period implied that the methanol solution should have been first saturated by hydrogen and equilibrium for the liquid phase/nitrogen flow should have been reached before the measurement became relevant to real gas evolution. As this process equally occurred for both photocatalysts, the difference in the initial  $H_2$  release could have been attributed to another reason: most probably to the transformation of the GO part of the parent photocatalyst *in situ* in the reaction mixture similarly to that had appeared previously under the hydrothermal pretreatment leading to HT-GO/TiO<sub>2</sub>.



## S References

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